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- Chapter 11 -

Stoichiometry Section

11.1 What is

stoichiometry? 1. true

2. true 3. false 4. true

5. true 6. 2, 2, 64.10 7.

3, 3, 96.00 8. 2, 2,

88.02 9. 4, 4, 72.08 10.

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methanol and oxygen  
gas 11. carbon dioxide  
and water 12. 160.10 g  
13. 160.10 g 14. They  
are equal. 15. A mole  
ratio is a ratio between  
the numbers of moles

### **VIBRATIONS AND WAVES**

Section 11.1 What is stoichiometry? In your textbook, read about stoichiometry and the balanced equation. For each statement below, write true or false. 2. 3.

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4. 5. The study of the quantitative relationships between the amounts of reactants used and the amounts of products formed by a chemical reaction is called stoichiometry.

Stoichiometry is based on the law of conservation of mass,

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Hydrocarbons ...

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13.1 Gases and Their

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describe the particle nature of both real and ideal gases. Read :  
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True/False Indicate whether the statement is true or false. \_\_\_\_ 1.

The actual yield is always lower than the theoretical yield.

Matching Match each item with the correct statement below. a. stoichiometry ...

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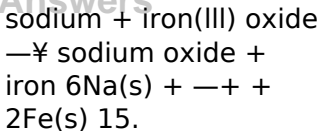
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In your textbook, read  
about mole ratios.

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Answer the questions about the following chemical reaction.



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Stoichiometry. As shown below, the correct mole ratio, 1 mol H<sub>2</sub> to 2 mol K, has moles of unknown in the numerator and

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moles of known in the denominator. Using this mole ratio converts the moles of potassium to the unknown number of moles of hydrogen.

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why reactions stop and how to determine the limiting reactant. Study the diagram showing a chemical reaction and the chemical equation that represents the reaction. Then complete the table.

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with a summary of

equation stoichiometry

problems and shows

how the skills

developed in Section

10.1 can be mixed with

the new skills

developed in this

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